

## Access Free Ch 14 Acids And Bases Practice Problems

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### **Ch 14 Acids And Bases**

Bronsted Lowry defines an acid as a proton donor/acceptor and a base as a proton donor/acceptor. amphoteric Substances--such as water--can act as acids or bases by either accepting or donating a proton.

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The cation from the base combines with the anion from the acid to make a salt. Acid is  $H^+$  DONOR, Base is  $H^+$  ACCEPTOR. In the reaction, the acid molecule gives an  $H^+$  to the base molecule. Acids the proton donor and bases the proton acceptors.

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Chapter 14 - Acids and Bases . 14.1 The Nature of Acids and Bases . A. Arrhenius Model 1. Acids produce hydrogen ions in aqueous solutions 2. Bases produce hydroxide ions in aqueous solutions B. Bronsted-Lowry Model 1. Acids are proton donors 2. Bases are proton acceptors 3.  $H_3O^+$  is called the hydronium ion C. Conjugate Acid- Base Pairs 1.

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(a) Fruits and fruit juices contain acids such as citric acid and ascorbic acid. Carbonated beverages contain benzoic acid, phosphoric acid, and carbonic acid. (b) Many household cleaners contain bases such as ammonia and sodium hydroxide. Antacids contain bases such as aluminum hydroxide.

## **CHAPTER 14 Acids and Bases - Weebly**

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Label the conjugate acid-base pairs in this system.  $\text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l})$  acid 1 base 2 base 1 acid 2 b. Is the forward or reverse reaction favored? Explain your answer. The forward reaction is favored. The weaker acid and weaker base are produced in the forward reaction.  $\text{HCO}_3^-$  competes more strongly with  $\text{H}_2\text{O}$  to donate a proton,

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MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

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## Chapter 14 - Acids and Bases | CourseNotes

4) In the Bronsted-Lowry model, a reaction between an acid and a base occurs to form a new acid and a new base. The acids and bases exist in pairs. When the acid loses a proton, it becomes a base and is called the conjugate base. When a base gains a proton, it becomes an acid and is called the conjugate acid.  $HA + B \rightleftharpoons A^- + BH^+$  (acid) (base) (conjugate) (conjugate) base acid

## Chapter 14: Acids and Bases

A conjugate base is formed when a Brønsted acid loses a proton. A conjugate acid is formed when a Brønsted base accepts a proton. Note: Water is amphoteric – it can act as an acid or a base depending on what its reacting with.

## Chapter 14: Creative Commons License Acids, Bases and

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## **Salts ...**

- strong acid: ionizes completely in aqueous solution.
- weak acid releases few hydrogen ions in aqueous solution.
- Strong Bases: Most bases are ionic compounds containing metal cations and the hydroxide anion,  $\text{OH}^-$ .
- $\text{NH}_3$  is molecular (not an ion)
- Ammonia produces hydroxide ions when it reacts with water molecules.

## **Acids and Bases: Chapter 14 & 15**

Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant ( $K_a$ ), & strong vs weak acids.

## **Chapter 14 (Acids and Bases) - Part 1**

CHAPTER 14 ACIDS AND BASES 501 24. A Lewis acid must have an empty orbital to accept an electron pair, and a Lewis base must have an unshared pair of electrons.

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## CHAPTER 14 ACIDS AND BASES

CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b.

## CHAPTER 14 REVIEW Acids and Bases

Major topics: polyprotic acids, salts as acids/bases, & salt solution pH calculations

## Chapter 14 (Acids and Bases) - Part 4

Chapter 14. Acid-Base Equilibria. 14.7 Acid-Base Titrations. Learning Objectives. By the end of this section, you will be able to: ... Substances such as phenolphthalein, which can be used to determine the pH of a solution, are called acid-base indicators. Acid-base indicators are either weak organic acids or weak

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organic bases.

## 14.7 Acid-Base Titrations - Chemistry

Arrhenius Acid-Base Theory Arrhenius Acid -a hydrogen containing compound that ionizes to produce hydrogen ions (  $H^+$  ) when dissolved in water Because molecular acids are not made of ions, they cannot dissociate.  $HCl(aq) \rightarrow H^+(aq) + Cl^-(aq)$  They must be pulled apart, or ionized, by the water.

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## Chapter 14: Acids and Bases

CHAPTER FOURTEEN ACIDS AND BASES Questions 16. The Arrhenius definitions are: acids produce  $H^+$  in water and bases produce  $OH^-$  in water. The difference between strong and weak acids and bases is the amount of  $H^+$  and  $OH^-$  produced. a. A strong acid is 100% dissociated in water. b. A strong base is 100% dissociated in water.



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## **CHAPTER FOURTEEN ACIDS AND BASES**

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AP Chemistry— CHAPTER 16 STUDY GUIDE— Acid-Base Equilibrium 16.1 Acids and Bases: A Brief Review •Acids taste sour and cause certain dyes to change color. •Bases taste bitter and feel soapy. •Arrhenius concept of acids and bases: •An acid is a substance that, when dissolved in water, increases the concentration of  $H^+$  ions.

### **AP Chemistry— CHAPTER 16 STUDY GUIDE Acid-Base Equilibrium**

Class 10 Science Chapter 2 Acids, Bases and Salts Notes - PDF

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