

## Chapter 9 Review Stoichiometry Section 1 Answers

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### Chapter 9 Review Stoichiometry Section

CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2. 200 g Given the ...

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. ... CHAPTER 9 REVIEW ...

### CHAPTER 9 REVIEW

Date.:SE(TIQf\$ | FCHAP] REV[EW. Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a masses in grams of all reactants and products.

### Date. FCHAP] REV[EW.

Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  CHAPTER 9 REVIEW ...

### CHAPTER 9 REVIEW

Chapter 9 - Stoichiometry Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a...

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CHAPTER 9 Stoichiometry Stoichiometry comes from the Greek ... SECTION REVIEW 1. What is stoichiometry? 2. How is a mole ratio from a reaction used in stoichiometric problems? 3. For each of the following chemical equations, write all possible mole ratios. a.  $2\text{HgO}(\text{s}) \rightarrow 2\text{Hg}(\text{l}) + \text{O}_2(\text{g})$

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Chapter 9 Section 2 covers Stoichiometric Calculations, including mole to mole, mole to mass, mass to mole, and mass to mass examples.

### 9.2 Ideal Stoichiometric Calculations

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1. \_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products.

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

### CHAPTER 9 REVIEW

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