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SECTION 2 PROBLEMS Write the answer

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on the line to the left. Show all your work in the space provided.

1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:

$$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$$

How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?

2. 200 g Given the ...

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and Chapter Summary ...

Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. ... CHAPTER 9 REVIEW ...

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CHAPTER 9 REVIEW

Stoichiometry. SECTION 2. PROBLEMS

Write the answer on the line to the left.

Show all your work in the space

provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$... CHAPTER 9 REVIEW ...

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CHAPTER 9 REVIEW

CHAPTER 9 REVIEW Stoichiometry

SECTION 2 PROBLEMS Write the answer on the line to the left Show all your work in the space provided

1 45 mol The following equation represents a laboratory preparation for oxygen gas:

$$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$$

How many

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moles of O_2 form if 30 mol

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This is the 1/2 half of the notes for
Chapter 9 over Reaction Stoichiometry.

Stoichiometry: Chapter 9

Chapter 10 155 Chapter 10 Chemical
Calculations and Chemical Equations
Review Skills 10.1 Equation

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10. 1.43 g/L calculate the volume of this gas. 76
STOICHIOMETRY MODERN CHEMISTRY a.
— . 81 g 6. A car air bag requires 70. L of
nitrogen gas ...

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SECTION 9.2. PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. The
following equation represents a

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laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many
grams of O_2 form if 3.0 mol of KClO_3 are
totally consumed? 2. Given the following
equation ...

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College Algebra, Chapter 9, Review

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Exercises, Section Review Exercises,
Problem 30 A teacher makes
\$35,000\$35,000 in his first year, and
gets a \$1,200\$1,200 raise each year. a.)
What is his salary A_n in his n th year
at this school? b.)

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MODERN CHEMISTRY SECTION 9-1
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work. Name Date Class CHAPTER 9
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SHORT ANSWER Answer the following
questions in the space provided. 1. The
coefficients in a chemical equation
represent the . (a) masses in grams of

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all reactants and products

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Chapter 9 Section 2 covers
Stoichiometric Calculations, including
mole to mole, mole to mass, mass to

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mole, and mass to mass examples.

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Chapter 9 Stoichiometry Definition •
Stoichiometry -Relationship between

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quantities • Composition stoichiometry
-The mass relationships of elements in
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chemical reaction Section 1 Introduction
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